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CHAPTER 11 REVIEW Gases

SECTION 3 SHORT ANSWER

Answer the following questions in

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the space provided. 1. _____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2.

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Three of the primary components of air are carbon dioxide, nitrogen,

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and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as $P_{\text{CO}_2} = 0.285$ torr and $P_{\text{N}_2} = 593.525$ torr. What is the partial pressure of oxygen?

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pressure of 0.200 atm to the
following units: 152 a. mm Hg 20.3
b. kPa 6. When an explosive like

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TNT is detonated, a mixture of gases at high temperature is created. Suppose that gas X has a pressure of 50 atm, gas Y has a pressure of 20 atm, and gas Z has a pressure of 10 atm. 80 atm a. What is the total ...

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CHAPTER 11 REVIEW Gases

SECTION 4 SHORT ANSWER

Answer the following questions in the space provided. 1. List the following gases in order of increasing rate of effusion.

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Chapter 11 Review Gases Section
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1 Answers Modern Chemistry 97
Gases CHAPTER 11 REVIEW
Gases SECTION 3 SHORT
ANSWER Answer the following

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Answers in the space provided. 1.

_____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided ...

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Quiz: Gas Volumes and the Ideal
Gas Law In the space provided,
write the letter of the term or
phrase that best completes each
sentence or best answers each
question. _____ 1. At the same

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Answer Key
temperature and pressure,
balloons of equal volume always
contain

Fundamentals of Chemistry,
Fourth Edition covers the

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fundamentals of chemistry. The book describes the formation of ionic and covalent bonds; the Lewis theory of bonding; resonance; and the shape of molecules. The book then discusses the theory and some applications of the four kinds of

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Answer Key: ultraviolet, infrared, nuclear (proton) magnetic resonance, and mass. Topics that combine environmental significance with descriptive chemistry, including atmospheric pollution from automobile exhaust; the metallurgy of iron and

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aluminum; corrosion; reactions involving ozone in the upper atmosphere; and the methods of controlling the pollution of air and water, are also considered.

Chemists and students taking courses related to chemistry and environmental chemistry will find

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the book invaluable.
Answer Key

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Worksheet 6: Liquids and Solids
MCQs Practice test Atomic
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to solve MCQ questions: Atoms,
atomic spectrum, atomic

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absorption spectrum, atomic emission spectrum, molecules, azimuthal quantum number, Bohr's model, Bohr's atomic model defects, charge to mass ratio of electron, discovery of electron, discovery of neutron, discovery of proton, dual nature of matter,

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Answer Key, electron charge, electron distribution, electron radius and energy derivation, electron velocity, electronic configuration of elements, energy of revolving electron, fundamental particles, Heisenberg's uncertainty principle, hydrogen spectrum, magnetic

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Answer Key
quantum number, mass of electron,
metallic crystals properties,
Moseley law, neutron properties,
orbital concept, photons wave
number, Planck's quantum theory,
properties of cathode rays,
properties of positive rays,
quantum numbers, quantum theory,

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higher ionization energies, ionic
radius, ionization energies,
ionization energy periodic table,
Lewis concept, and modern
periodic table. Practice test
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Answer Key chromatography,
crystallization, filter paper
filtration, filtration crucibles,
solvent extraction, and
sublimation. Practice test Gases
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MCQ questions: Gas laws, gas
properties, kinetic molecular

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theory of gases, ideal gas constant, ideal gas density, liquefaction of gases, absolute zero derivation, applications of Daltons law, Avogadro's law, Boyle's law, Charles law, Daltons law, diffusion and effusion, Graham's law of diffusion, ideality deviations,

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boiling points, crystal lattice,
crystals and classification, cubic
close packing, diamond structure,
dipole-dipole forces, dipole
induced dipole forces, dynamic
equilibrium, energy changes,
intermolecular attractions,
hexagonal close packing, hydrogen

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Answer Key
bonding, intermolecular forces,
London dispersion forces, metallic
crystals properties, metallic solids,
metal's structure, molecular solids,
phase changes energies,
properties of covalent crystals,
solid iodine structure, unit cell, and
vapor pressure.

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various possible interpretations of structures, reactions and spectra. Chemical reasoning based on simpler quantum models is, however, essential to enumerating the likely possibilities. The simpler models also often suggest the type of wave function likely to be

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involved in ground and excited states at various points along reaction paths. This preliminary understanding is needed in order to select the appropriate higher level approach since most higher level models are designed to describe improvements to some

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reasonable zeroth order wave function. Consequently, most of the chapters in this volume begin with experimental facts and model functions and then progress to higher level theory only when quantitative results are required. In the first chapter, Zimmerman

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discusses a wide variety of thermal and photochemical reactions of organic molecules. Gronert discusses the use of ab initio calculations and experimental facts in deciphering the mechanism of β -elimination reactions in the gas phase. Bettinger et al focus on

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considerations involved with many of the methods and illustrates the difficulties by comparing with the UV spectra of short polyenes. Jordan et al discuss long-range electron transfer using model compounds and model Hamiltonians. Finally, Hiberty

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discusses the breathing orbital valence bond model as a different approach to introducing the crucial correlation that is known to be important in organic reactions. Contents: Some Theoretical Applications to Organic Chemistry (H E Zimmerman) Ab Initio Studies

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Gronert) Computational Analyses

of Prototype Carbene Structures

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Organic Diradicals — Where to

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Identify Them (D A Hrovat & W T Borden) Ab Initio Methods for the Description of Electronically Excited States: Survey of Methods and Selected Results (R J Cave) Long-Range Intramolecular Interactions: Implications for Electron Transfer (K D Jordan et

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al.) The Breathing Orbital Valence
Bond Method (P C Hiberty)

Readership: Graduate and
postgraduate students in organic
chemistry. keywords: Electronic
Structure; Organic
Chemistry; Reaction
Mechanisms; Carbene

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Structures;Diradicals;Excited
States;Long-Range
Interaction;Valence Bond
Theory;Molecular Orbitals

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In this new textbook on physical chemistry, fundamentals are introduced simply yet in more depth than is common. Topics are arranged in a progressive pattern, with simpler theory early and more complicated theory later.

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General principles are induced from key experimental results. Some mathematical background is supplied where it would be helpful. Each chapter includes worked-out examples and numerous references. Extensive problems, review, and discussion questions

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Answers Key are included for each chapter.

More detail than is common is devoted to the nature of work and heat and how they differ.

Introductory Caratheodory theory and the standard integrating factor for dG_{rev} are carefully developed. The fundamental role played by

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Uncertainty and symmetry in quantum mechanics is emphasized. In chemical kinetics, various methods for determined rate laws are presented. The key mechanisms are detailed. Considerable statistical mechanics and reaction rate theory are then

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Answer Key bonds, and condensed states of matter. This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists.

Separation and purification are covered in the first chapter, while

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the following chapters focus on atomic and molecular weights, stoichiometry, the structure of atoms, and types of chemical bonds. The molecular orbital (MO) theory of bonding, galvanic cells, and chemical thermodynamics are considered next. Separate

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chapters are devoted to MO theory of covalent and metallic bonding; orbital hybridization; intermolecular forces; acids and bases; ionic equilibrium calculations; and polymers and biochemicals. This monograph is intended for chemistry students.

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