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Answers
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 $= \text{NOBr(g)}$ $K_p = 5.3$ $2\text{NO(g)} = \text{N}_2\text{(g)} +$
 $\text{O}_2 \text{(g)}$ $K_c = 2.1 \times 10^3$ Use these
reactions and their equilibrium
constants to predict the equilibrium
constant for the following reaction:
 $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} + \text{Br}_2 \text{(g)} = 2\text{NOBr(g)}$
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