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Valence Electrons
and the Periodic
Table
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Electron Behavior. ...

-ordered arrangement
by wavelength or
frequency for all forms
of electromagnetic
radiation. Parts of the
wave. Wavelength-
 λ () The
distance between
corresponding points

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4 on adjacent waves.
Units: m, nm, cm, or Å

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Electrons Chapter 4
CHAPTER 4 REVIEW
Arrangement of
Electrons in Atoms
SECTION 3 SHORT
ANSWER Answer the
following questions in
the space provided. 1.
State the Pauli
exclusion principle,

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4 and use it to explain why electrons in the same orbital must have opposite spin states. The Pauli exclusion principle states that no two electrons in an atom may have the

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Answers

Pacing Regular

Schedule: with lab(s):

3 days without lab(s):

2 days Block

Schedule: with lab(s):

1 1/2 days without

lab(s): 1 day

Objectives 1. Explain

the mathematical

relationship between

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4 and 5, which deals with electrons and the periodic table.

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Periodic Law] Chapter
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Bonding] ...
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Vocabulary:

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to Know for the Final
From This Chapter ...
the arrangement of
electrons in an atom.

Ground-State

Electron

Configuration. the

lowest-energy

arrangement of the

electrons for each

element. Aufbau

Principle.

Chapter 4

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Elements & Electron
Configurations

Elements of the 6th
and 7th periods
contain $14f$ orbitals. Do
...

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Chapter 4:

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Electromagnetic
Radiation.

Electromagnetic
Spectrum.

Wavelength.

Frequency. The
radiation associated
within electric and
magnetic field; ν . All
of the frequencies or
wavelengths of

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CHAPTER 4 REVIEW

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SECTION 3 SHORT

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ANSWER Answer the following questions in the space provided. 1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states.

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Arrangement of
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Chapter 4: Arrangement

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Section 4-3: Electron Configurations _____

Pacing Regular

Schedule: with lab(s):

NA without lab(s): 2

days Block Schedule:

with lab(s): NA

without lab(s): 1 day

Objectives 1. List the

total number of

electrons needed to

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fully occupy each
main energy level. 2.

Chapter 4:

Arrangement of
Electrons in Atoms

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ARRANGEMENT OF
ELECTRONS IN

ATOMS The following
pages contain the
bulk (but not all) of the
information for the
chapter 4 test. Focus

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on this content, but
make sure to review
class notes, activities,
handouts, questions,
etc. If you study this
document and
NOTHING else, you
should at least be
able to PASS the test.

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Chapter 4
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Electrons orbitals of

equal energy are
each occupied by one
electron before any
orbital is occupied by
a second electron,
and all electrons in
singly occupied
orbitals must have the
same spin Pauli's
exclusion Page 5/26

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Chapter 4
Arrangement Of
Electrons - Bespokify

The Pauli exclusion
principle states that
no two electrons in an
atom may have the

Chapter Four
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CHEMISTRY

CHAPTER 4.

(Arrangement of

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4(Arrangement) The lowest energy state of an atom is its ground state.

Answers

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How was Rutherford's model incomplete Did not explain how the negatively charged electrons distributed itself in the electron cloud around the positively charged nucleus without being attracted to each other Holt Modern Chemistry: Chapter 4 Flashcards | Quizlet

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