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~~Acids and Bases~~

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~~Introduction Acids~~

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Acids Bases) - Part 4

*Chapter 14 (Acids
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*Chapter 14 (Acids
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Ka Kb Kw pH pOH

pKa pKb H+ OH-

Calculations - Acids

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~~Chapter 16 Acid-Base Equilibria~~

pH, pOH, H_3O^+ ,
OH⁻, K_w , K_a , K_b ,
pKa, and pKb Basic
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Chapter 14 Acids

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14 - Acids and

Bases. 14.1 The

Nature of Acids and

Bases. A. Arrhenius

Model 1. Acids

produce hydrogen

ions in aqueous

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solutions 2. Bases produce hydroxide ions in aqueous solutions B.

Bronsted-Lowry

Model 1. Acids are proton donors 2.

Bases are proton acceptors 3.

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Acidic*acid* ** low*

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Introduction.

Originally the terms acid and base referred to taste. The practice of classifying substances according to their acidic (sour) or basic (alkaline or bitter) properties dates back to ancient times. An acid was

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Chapter 14

something with a sour taste, such as lemon juice, and a base was something with a bitter taste, such as tonic water.

*Introduction to
Acids and Bases
(Worksheet) -
Chemistry ...
Acid and Base
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Answers. 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$. HNO_3 and NO_3^- make one

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Chapter 14

pair OH^- and H_2O
make the other. b)
 $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightleftharpoons$
 $\text{CH}_3\text{NH}_3^+ + \text{OH}^-$

*Acid and Base
Worksheet -*

Answers -

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Chapter 14 Acids
And Bases an acid
that is composed of
H, O, and a third

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element. The H & O are covalently bonded and are called the hydroxyl group. Property of bases. taste bitter, feel slippery, change color of indicators, react with acids to produce a salt and water neutralization rxn, are electrolytes.

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Arrhenius or
traditional base.

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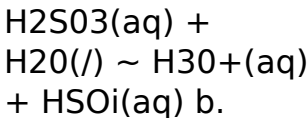
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REVIEW . Acids and
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ANSWER Answer
the following
questions in the
space provided. 1.

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a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1:



b. Which stage of ionization usually produces more ions? Explain your answer.

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