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Stoichiometry,

Theoretical, Actual &

Percent yield Part I.

Stoichiometry 1. 1

$ZnI_2 \rightarrow 1 Zn + 1 I_2$

How many grams of

iodine will you

produce if you begin

your rxn with 56.7g

of zinc iodide? 56.7 g

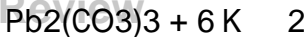
1 mole 1 mole 253.8

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g 319.19 g 1 mole 1

mole = 45.08 grams
of Iodine 2. 1



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Stoichiometry,

Theoretical, Actual ...

STOICHIOMETRY 12

Vocabulary Review

Match the correct

vocabulary term to

each numbered

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statement. Write the letter Of the correct term on. the line.

Column A 1. the starting materials in a chemical reaction 2. a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms Of moles

Mister Chemistry

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Stoichiometry.

STUDY. PLAY. Mole.

6.02×10^{23}

representative

particles. Reactants. A

starting material in a

chemical reaction.

Products. The

elements or

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compounds
produced by a
chemical reaction.
Limiting reactant.

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Important Vocabulary •

Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry

- Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted

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in terms of moles
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Stoichiometry. You react 25.0 liters of oxygen gas at STP with 50.0 ml of ethanol liquid.

Ethanol has a density of 0.789 g/ml. Use this reaction to solve the following problems.

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