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Answers

Baking Soda and
Vinegar
Stoichiometry Lab
Experiment ~~Acetic~~
~~Acid \u0026amp; Baking~~
~~Soda Stoichiometry~~
~~Lab: Calculating~~
~~theoretical yield of~~
~~CO₂ Air Bag~~

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~~Stoichiometry Lab~~
Lab: Where Did it
Go? Stoichiometry
of a Household
Reaction

Stoichiometry
& Law of
Conservation of
Mass
Vinegar and
Baking Soda
Stoichiometry Lab
Stoichiometry
Chemistry Lab -
Decomposition of

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Baking Soda

~~Baking Soda Air~~

~~Bag Lab I~~

~~Chemistry Matters~~

~~Quitalig, Joanna~~

~~Laboratory~~

~~exercise~~

~~Stoichiometry 1~~

Target

Stoichiometry Lab

Acetic acid and

baking soda for

Limiting Reactants

Backyard

Chemistry -

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Stoichiometry with
baking soda and
vinegar Pre-lab
decomposition of

baking soda

~~Limiting Reactant~~

~~Demonstration~~

How to Convert

Baking Soda into

Washing Soda

Chemical Reaction

Of Baking Soda And

Vinegar (Sodium

Bicarbonate And

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Acetic Acid)

Stoichiometry
CHEM111L:

Lab Report
Bicarbonate

Answers
Decomposition

Post-Lab Video

Stoichiometry lab

Na₂CO₃ to NaCl

Stoichiometry

Decomposition of
sodium

bicarbonate Lab

~~Limiting Reagents~~

~~Lab video~~ Lab

Experiment 8 -

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Baking Soda

Vinegar Air Bags

Lab #9 - Mole
Ratios and
Reaction

Stoichiometry

~~STOICHIOMETRY~~

~~LABORATORY 001~~

Baking Soda Lab -
Percent Yield Chem
10 Reaction

Stoichiometry Lab

Decomposition of

Sodium

Bicarbonate

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Baking Soda

~~(Baking Soda) Lab~~

~~1.1 Heating Baking
Soda The~~

~~chemistry of~~

~~cookies Stephanie~~

~~Warren Limiting~~

~~Reagent Lecture~~

~~Decomposition of~~

~~Baking Soda Baking~~

~~Soda Stoichiometry~~

~~Lab Report~~

This lab

demonstrates the

reactivity of two

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Baking Soda

household cooking items, baking soda and vinegar.

Baking soda is a powdered chemical compound called sodium bicarbonate, and vinegar includes acetic acid. These 2 components react in solution to form carbon dioxide, water, and

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Sodium acetate as
shown in the
chemical reaction
below:

~~Stoichiometry:~~

~~Baking Soda and~~

~~Vinegar Reactions~~

Lab 21:

Stoichiometry –

Decomposition of

Baking Soda Safety

☐ Handle the

contents from

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Stoichiometry
Lab Report
Answers

Stove with care to prevent burns. Pre-Lab Overview:

Have you ever baked? Baking soda (sodium bicarbonate, NaHCO_3) is used in bakery products to ensure that they rise during baking. Why? As the dough is heated, the baking soda

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decomposes,

Stoichiometry

Lab Report

~~Lab 21:~~

~~Stoichiometry—~~

~~Decomposition of
Baking Soda~~

There are three
theoretically
possible chemical
reactions that
could occur during
the thermal
decomposition of
baking soda. 1)

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Baking Soda

Sodium

bicarbonate (s) →

sodium hydroxide

(s) + carbon

dioxide (g) 2)

sodium

bicarbonate (s) →

sodium oxide (s) +

carbon dioxide (g)

+ water (g) 3)

sodium

bicarbonate (s) →

sodium carbonate

(s) + carbon

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Baking Soda

dioxide (g) + water
(g)

Lab Report

~~Answers~~
~~Lab Report~~

~~Stoichiometry~~

~~Decomposition of~~

~~Sodium ...~~

1. Find the mass of the evaporating dish and watch glass. Record this mass in the Data Table. 2. Add 1/3 of a teaspoon of

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Lab Report
Answers

baking soda to the evaporating dish, and record the total mass in the Data Table. 3.

Cover the evaporating dish with the watch glass so that only the spout of the evaporating dish is exposed.

~~Stoichiometry and~~

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~~Baking Soda Lab~~

Data &
Lab Report

Observations DATA

ANSWERS
TABLE Actual

amount of sodium
hydrogen

carbonate (baking
soda/ NaHCO_3)

used: 4.2 g

Expected
(calculated)

amount of sodium
acetate to be

produced: 4.1 g

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Mass of empty 500-
mL flask: 108.9.

Mass of 500mL
flask after water
has evaporated:

112.1 Actual mass
of sodium acetate
produced: 3.2

Percent Yield of
Sodium Acetate

produced: $3.2 \div 4.1$
 $\times 100 = 0.78 \times 100$
 $= 78.0$

OBSERVATION

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TABLE Three
physical properties
of sodium
hydrogen
carbonate (baking
soda/ NaHCO_3):
powdery ...

~~Stoichiometry Lab~~
~~Report - Weebly~~
Stoichiometry Lab
Report Brittney
Aceron Karla Wade-
Choza, Jonathan

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Guerrero, Luis

Martinez...
Lab Report

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Docs~~

Lab Hints □

Students may ask how much of the baking soda they should use. In keeping with the general practice of not filling a crucible

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more than half-full, there is no "correct" mass of baking soda to use. This avoids situations where students believe they must use 2.00 g of baking soda or else the experiment "won't work."

~~Decomposition of~~

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~~Baking Soda—Flinn
Stoichiometry
Scientific~~

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Answers~~

In this particular lab we used stoichiometry, the part of chemistry that studies amounts of substances that are involved in reactions, to observe the reactions made by combining sodium

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Docs~~

On the second day they conduct the lab, and on the third day they write and critique their lab report. In this lesson students learn how to design an experiment in

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which they can evaluate how closely an experiment's actual yield corresponds to the theoretical yield. For the hypothesis, students use stoichiometry to predict how much carbon dioxide is produced when mixing a known

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amount of vinegar
and baking soda.

~~Answers~~
Eleventh grade

~~Lesson~~

~~Stoichiometry~~

~~Experimental~~

~~Design~~

Part A: Baking soda

(NaHCO_3) and

vinegar ($\text{C}_2\text{H}_4\text{O}_2$)

in a closed

Ziploc bag 1.

Safety glasses

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were put on 2. Ziploc bags were labelled "Ziploc bag 1" and "Ziploc bag 2" 3. 10ml of baking soda was measured into a small beaker. 4. The measured 10ml of baking soda was poured into Ziploc Bag 1. 5. 15ml of vinegar was measured into

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a...
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Investigation into
Answers
Conservation of
Mass Lab~~

This lesson is part of a three-day lab. In the first day students design their lab, which includes solving a stoichiometry problem. On the second day they

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conduct the lab, and on the third day they write and critique their lab report. In this lesson students will conduct a lab that they planned in the previous lesson. In their experimental design, students used stoichiometry to predict how much carbon

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dioxide would be produced from a set amount of vinegar and baking soda.

~~Stoichiometry Lab Report~~

~~BetterLesson~~

Name _Aisha

Wint _____ Date

_____ Jan 11,

2020 _____

Experiment -

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Stoichiometry

Determining the
Limiting Reagent

Using the Reaction
of Sodium

Bicarbonate with
Acetic Acid

Materials provided
in the kit: 100 mL
graduated cylinder

Materials provided
by the student:

Four sandwich size
zip-lock bags

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Stoichiometry

(sodium
Lab Report

bicarbonate,

Answers
NaHCO₃) White

vinegar Container

in which to set your

bags during the ...

~~A_Wint__Lab_Limiti~~

~~ng_Reagent.docx~~

~~Name_Aisha Wint~~

~~Date ...~~

Procedure Our

Ourbag Experiment

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Objective In order to create out air bag we need a Ziploc bag, baking soda, vinegar, a mini plastic bag, a rubber band and tape. First we poured 200ml of vinegar into a beaker and poured it into the Ziploc back. We then took the mini plastic bag

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Stoichiometry

~~Airbag Lab by~~
~~Sabrina Wright~~
Prezi

Pre-lab discussion:

1. How is our lab experiment similar to a real airbag's reaction and how is it different? 2.

Summarize the objective of the lab. Background:

You will use

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Stoichiometry

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Answers

quantities of baking soda and vinegar to maximize the amount of CO_2 gas created and minimize added mass due to unreacted vinegar or baking soda.

~~Stoichiometry Air Bag Lab~~

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~~Stoichiometry~~

In this lab, you will need to do a

reaction where

baking soda will

react with an.

Aspirin is also

present in Alka-

Seltzer tablets to

reduce fever and

relieve headaches,

but in this lab, we

are going to study

the reaction that

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Baking Soda

Stoichiometry

takes place
between. Report
Sheet for

Answers
Stoichiometry Lab:
Reaction of Sodium
Bicarbonate with
Acetic.

~~Stoichiometry lab
report | Spectrum~~

Vinegar and Baking
Soda Stoichiometry
Lab Purpose: To
predict the amount

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of Carbon Dioxide
gas that should be
produced in a
chemical reaction;

then calculate the
amount of CO₂
released, the
percent yield. 00

Grams of a
Compound?

Student Laboratory
Worksheet,
continued 5. A
standardized

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Solution is a
solution of known
molarity.

Lab Report
Answers

~~Stoichiometry lab
experiment
answers~~

~~CDiscount~~

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Lab 11 Report from
CHEM 3571, 3572
at Gaithersburg
High.

Stoichiometric

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Determinations Lab

Stoichiometry and

Lab Report
Limiting Reactant

Answers
PURPOSE To find

the limiting

reactant and

measure Eleventh

grade Lesson

Stoichiometry

Experimental

Design KEY

Chemistry:

Stoichiometry and

Baking Soda

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(NaHCO_3)
Stoichiometry
Lab Report

Purposes: 1.

~~Answers~~
~~Stoichiometric 11~~
~~Determinations Lab~~
~~Answers~~

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paper. Apply stoichiometry and the idea of a limiting reactant to a reaction in solution. In this challenge you will test your stoichiometric prowess in answer to the.

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